



## Adsorption of Phenol from Refinery Wastewater Using Rice Husk Activated Carbon

<sup>1</sup>Y.S. Mohammad, <sup>2</sup>E.M. Shaibu-Imodagbe, <sup>1</sup>S.B. Igboro, <sup>3</sup>A. Giwa and <sup>1</sup>C.A. Okuofu

<sup>1</sup>Department of Water Resources and Environmental Engineering, A. B. U. Zaria, Nigeria

<sup>2</sup>Samaru College of Agriculture, Division of Agricultural Colleges, A. B. U. Zaria, Nigeria

<sup>3</sup>Department of Textile Science and Technology, A. B. U. Zaria, Nigeria

Received: August 8, 2014; Accepted in Revised form: October 5, 2014

**Abstract:** Industries such as petroleum refineries continuously generate large volumes of wastewater that contain high concentrations of phenol; therefore wastewater treatment as an integral part of their activities is required. The main objective of this paper is to demonstrate the applicability of rice husk activated carbon in an adsorption column for the treatment of phenolic refinery wastewater. Effects of bed depth (of the obtained rice husk activated carbon) and flow rate (of the waste water) on the sorption of phenol from the waste water were investigated. Column models such as Bed Depth Service Time, Thomas and Yoon Nelson models were employed to analyze the experimental data. The column adsorption experiment attained adsorption capacity of 28mg/g at breakthrough point of 0.5, flow rate of 4.5ml/min and bed depth of 7.5cm. The performance of the column adsorption was affected by bed depth and flow rate and models such as Thomas and Yoon Nelson could be employed in estimation of design parameters and scale-up operations during pilot or full scale application.

**Key words:** Rice husk • Adsorption model • Adsorption column • Refinery wastewater

### INTRODUCTION

Adsorption is a process which involves a solid phase (sorbent material) and a liquid phase (solvent, normally water) containing a dissolved species to be sorbed (sorbate, metal ions). Due to higher affinity of the sorbent for sorbate species, the latter is attracted and bound there by different mechanisms. The process continues till equilibrium is established between the amount of solid-bound sorbate species and its portion remaining in the solutions. The degree of sorbent affinity for the sorbate determines its distribution between the solid and the liquid phases [1].

The discharge of industrial wastewaters with high concentrations of chemical contaminants results in chemical hazards on the receiving environmental component [2]. These hazards include injury and damage to environmental resources among which are aesthetic nuisance, health disorder, damage to aquatic life and degradation of environmental resources [3-5]. Phenols are among the most common water pollutants that can cause

hazards including health hazards which mostly results in death of living organisms including man. Processing industries such as petroleum refineries continuously generate large volumes of wastewaters that contain high concentration of phenols and other hydrocarbons. These phenolic compounds are among the main dominant group of compounds in refinery waste waters and this group has also been reported to be one of the difficult hydrocarbons to degrade biologically [6, 7]. The enhancement of the environmental friendliness and sustainability of the processing industries entails wastewater treatment to remove these phenolic compounds as an integral part of these processing activities of these plants. This will minimize the environmental damage from the discharge of the resulting wastewaters from the processing industries including the petroleum refineries. Ultimately, the cost of such treatment adds to the overall cost of production. From economic and environmental considerations, this necessitates the continuous search for simple, cleaner and cost-effective treatment option.

**Corresponding Author:** Y.S. Mohammad, Department of Water Resources and Environmental Engineering, A. B. U. Zaria, Nigeria. Tel: +2348035744222, E-mail: yahsaymoh@yahoo.co.uk.

Most of the treatment methods for wastewaters are not without their drawbacks such as high capital and operational cost, regeneration cost and residue disposal [8-10]. Consequently, biosorption has been identified as one of the most efficient technique for the removal of wastewater contaminants because of the potential for low-cost adsorbent especially activated carbon from agricultural wastes [11-15]. Hence there is increased attention on utilization of agricultural wastes in adsorption processes [37]. More increasing demands for food production will lead to additional generation of agricultural wastes such as rice husk which increase additional challenge in solid waste disposal. For example, in an attempt to become self-sufficient in rice production, Nigeria has been on the path of boosting its rice production which will invariably lead to increased generation of rice husk (as solid waste) in the environment [16]. Undoubtedly, the fulfilment of increased rice production will generate additional environmental challenge in sustainable solid waste management of the resulting rice husk. There is little or no information on the application of rice husk derived activated carbon in the biosorption of phenolic components using column adsorption from refinery wastewaters. Therefore, the objective of this paper is to highlight the applicability of rice husk activated carbon in the biosorption of phenolic components from refinery wastewaters using a column adsorption arrangement.

## MATERIALS AND METHOD

**Laboratory Materials/Reagents:** Furnace, ceramic flat surface, oven drier, de-ionized water, measuring cylinder, digital weighing balance, standard flask, conical flask, retort stand and clamp, micro filter, syringe, centrifuge, UV spectrophotometer, quartz cuvettes, glass column, glass wool, connecting pipe and valve,  $H_3PO_4$ , and phenol.

**Production of Activated Carbon:** The natural precursor used in the preparation of adsorbent is rice husk which was collected from National Cereal Research Institute, Badeggi. When collected from NCRI Rice Mill, rice husk was washed with distilled water to remove dirt and surface impurity, then oven-dried at  $100^\circ C$  for 24h [17]. In the thermal pretreatment, rice husk was placed on a ceramic flat surface, charged into a furnace and heated to a temperature of  $441.46^\circ C$  (optimum condition developed in preliminary studies) at a heating rate of  $20-25^\circ C/min$  and residence time of 1h. The charred residue was collected and cooled at room temperature. In the chemical pretreatment, the carbonized rice husk (charred residue)

was activated with  $1M H_3PO_4$  for 3h at impregnation ratio of 2:1 (volume *ml* of acid/mass *g* of rice husk) and later oven-dried overnight at  $200^\circ C$  to ensure proper drying [18, 19]. The material was then removed from the oven, cooled for 2h and then washed with distilled water to bring the pH to 7.0 and again oven-dried overnight at  $100^\circ C$  [19].

**Column Experiment:** The optimum operating condition developed in our previous study was applied in the column experiment. That is rice husk activated carbon was produced at optimum temperature of  $441.46^\circ C$  and the optimum dosage of 4g was the basis for the establishment of the bed height/depth of 6cm. The column experiment was conducted using a glass column of 12cm length, 2cm internal diameter,  $3.142cm^2$  surface area and  $37.7cm^3$  empty bed volume. The column was packed with a known quantity of rice husk activated carbon to yield the desired bed depth/height. Glass wool was placed at the bottom as a support and also at the top to prevent flotation of rice husk activated carbon in excess wastewater [20]. Process wastewater, which is the wastewater generated from various unit processes/operations, was collected from Kaduna Refinery and Petrochemical Company before treatment. Phenol composition of the process wastewater was analyzed and found to be  $51.32mg/l$ . This initial concentration was used throughout the experiment. The wastewater was fed to the column in a downward flow mode through a pipe having a valve as the flow regulator. The effluent wastewater was collected from the column outlet at predetermined time interval for analysis of unadsorbed phenol concentration.

The effect of the following variables on the column performance was investigated:

-Bed depth: 7.5cm (5g), 6cm (4g), 4.5cm (3g); keeping the flow rate constant at 3ml/min

-Flow rate: 3ml/min, 4.5ml/min; keeping the bed depth constant at 6cm (4g).

**Breakthrough Curve:** The breakthrough curves for different bed depths and flow rates were generated by plotting  $C_B/C_0$  against  $t$ . The breakthrough characteristics for each point were thereafter obtained from the curves.

**Breakthrough Capacity:**  $q_B$  Breakthrough capacity was determined using the equation [21]:

$$q_B = \frac{t_B f C_0}{m} \quad (1)$$

**Volume of Effluent Treated  $V_{eff}$ :** Volume of effluent treated was determined using the equation [22]:

$$V_{eff} = t_B f \quad (2)$$

**Removal Efficiency:** Removal efficiency was determined using the equation [23-25]:

$$RE = \frac{(C_o - C_B)}{C_o} \times 100 \quad (3)$$

where:  $q_B$  = breakthrough adsorption capacity (mg/g),  $V_{eff}$  = volume of effluent (L),  $C_B$ ,  $C_t$  = effluent concentration of solute at breakthrough or time  $t$  (mg/l),  $t_B$  or  $t$  = breakthrough time or time (min),  $C_o$  = influent concentration of solute (mg/l),  $f$  = volumetric flow rate (l/min),  $C_e$  = equilibrium concentration of solute (mg/l),  $m$  = mass of activated carbon used (g),  $RE$  = removal efficiency (%).

**Column Adsorption Models:** For the column adsorption study, the column models below were used to analyze the experimental data.

**Bed Depth Service Time (Adam Bohart) Model:** The Bed Depth Service Time (BDST) model equation given below was used by plotting service time  $t$  against bed depth  $z$ . Model parameters were obtained from the slope and intercept of the plot [26, 27]:

$$t = \frac{N_o}{C f} Z - \frac{1}{C_o K_a} \ln \left( \frac{C_o}{C_B} - 1 \right) \quad (4)$$

The equation above can be written as [26, 28]:

$$t = aZ - b \quad (5)$$

where:

$$a = slope = \frac{N_o}{C_o F}, b = intercept = \frac{1}{C_o K_a} \ln \left( \frac{C_o}{C_B} - 1 \right)$$

**Thomas Model:** The Thomas model equation given below was used to analyze the column experimental data by plotting  $\ln \frac{C_o}{C_B} - 1$  against  $t$ . Model parameters were obtained from the slope and intercept of the plot [29]:

$$\ln \left( \frac{C_o}{C_B} - 1 \right) = \frac{K_{TN} q_o m}{f} - k_{TN} C_o t \quad (6)$$

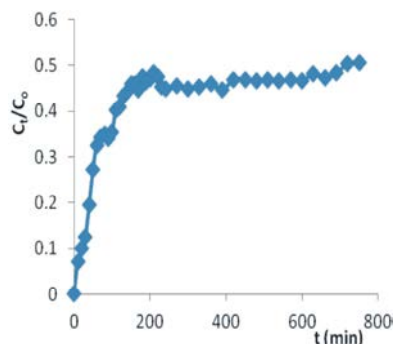


Fig. 1: Breakthrough curve for Bed depth, 6cm (4g) and Flow rate, 4.5ml/min

**Yoon Nelson Model:** The Yoon Nelson model equation given below was also used to analyze the column experimental data by plotting  $\ln \frac{C_o}{C_o - C_B}$  against  $t$ . Model parameters were obtained from the slope and intercept of the plot [9]:

$$\ln \left( \frac{C_o}{C_o - C_B} \right) = k_{YN} t - k_{YN} \tau \quad (7)$$

where:  $K_a$  = adsorption rate constant (l /mg. min),  $F$  = linear flow velocity to bed (m/min),  $C_B$  = effluent concentration of solute at breakthrough (mg/l),  $C_o$  = influent concentration of solute (mg/l),  $N_o$  = adsorption capacity (mg/l),  $Z$  = bed depth (m),  $t$  = time (min),  $K_{th}$  = Thomas rate constant (l/(min mg)),  $f$  = volumetric flow rate (l/min),  $q_o$  = adsorption capacity (mg/g),  $m$  = mass of the adsorbent (g),  $K_{YN}$  = Yoon Nelson rate constant,  $\tau$  = breakthrough time at 50% (min)

## RESULTS AND DISCUSSION

**Column Adsorption Capacity and Removal Efficiency:**

It was observed that as the column adsorption proceeds with time, the adsorption capacity increases while the removal efficiency decreases. This is because as the influent stream pass through the bed, more solute molecule are captured and retained by the activated carbon and this continuously increases the adsorption capacity until the bed becomes saturated. But as this continues with time, the amount of solute captured from the stream and retained by the activated carbon decreases and this result in an increase in the amount of solute concentration in the effluent stream, hence removal efficiency will be observed to be decreasing. It was deduced from Fig. 1 that at breakthrough point of 0.5, the bed attained adsorption capacity of 28mg/g.

Table 1: Break-through Characteristics

Breakthrough points (b.p)			0.1			0.2			0.4	
Z (cm)	f (ml/min)	t <sub>B</sub> (min)	V <sub>eff</sub> (ml)	q <sub>B</sub> (mg/g)	t <sub>B</sub> (min)	V <sub>eff</sub> (ml)	q <sub>B</sub> (mg/g)	t <sub>B</sub> (min)	V <sub>eff</sub> (ml)	q <sub>B</sub> (mg/g)
4.5	3	23	69	1.180	54	162	2.771	97	291	4.978
6	3	44	132	1.694	77	231	2.974	165	495	6.351
7.5	3	69	207	2.125	103	309	3.172	290	870	8.930
6	4.5	21	94.5	1.212	42	189	2.425	110	484	6.210

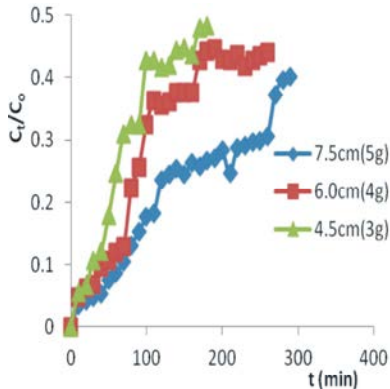


Fig. 2: Breakthrough curve for different Bed depth

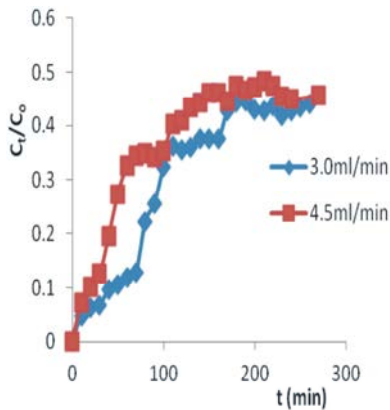


Fig. 3: Breakthrough curve for different Flow rate

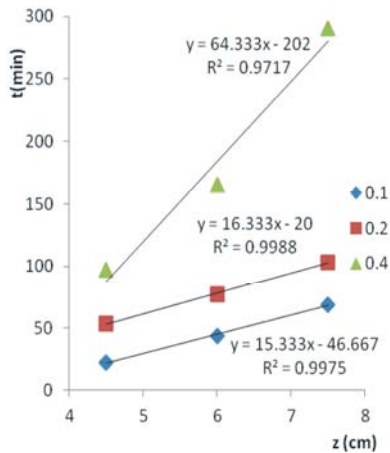


Fig. 4: BDST model plot for different Breakthrough points

**Effect of Process Variables**

**Effect of Bed Depth:** From Fig. 2 it could be observed that the breakthrough time increases with increase in bed depth. This observation is similar to that of Zhang *et al.*, [30]; Ji *et al.*, [31] and it is probably due to increased surface area and availability of more binding sites for sorption as the bed depth increases [32]; thereby resulting in a broadened mass transfer zone. This means that as the bed depth is increased the length of the bed which the wastewater passes through also increases. Table 1 shows the breakthrough characteristics with respect to increase in bed depth. From the table also, at any breakthrough point, the corresponding breakthrough time, capacity and volume of wastewater treated increases with increase in bed depth. It can also be observed from the table that the larger the breakthrough time, the higher the breakthrough capacity of the column. This could be as a result of the fact that increase in bed depth provides the solute (phenol) molecule with more time to get adsorbed into the broadened mass transfer zone and this also result in treating more volume of effluent [22, 27]. Therefore, a decrease in the phenol composition of the effluent can be achieved in a column design approach by increasing the bed depth at the same time [27, 33]. Hence, higher bed depth favours better column performance.

**Effect of Flow Rate:** Fig. 3 shows that the breakthrough time decreases with increase in flow rate. This observation is in agreement with the trend reported by Zhang *et al.*, [30] and Kundu *et al.*, [34]. Table 1 also shows the breakthrough characteristics with respect to increase in flow rate. The table also shows that relative breakthrough time, capacity and volume of waste water at any breakthrough point decreases as the flow rate increases. This is because as the flow rate is increased, more volume of the wastewater is made to pass through the adsorption zone thereby providing insufficient residence time for the solute molecule to adequately transverse to the functional group-containing pores [22]. This leads to the exit of solute in the wastewater from the column before the establishment of equilibrium [35] and usually results in shorter breakthrough time and capacity. Hence, lower flow rate favours better column performance.

Table 2: Bed Depth Service Time (BDST) model parameters

b.p	$C_0/C_1 - 1$	$\ln(C_0/C_1 - 1)$	a	b	$N_0$ (mg/l)	$K_a$	$R^2$
0.1	9	2.197	15.333	46.667	751.122	0.000917	0.9975
0.2	5	1.609	16.333	20	800.109	0.001568	0.9988
0.4	2.5	0.916	64.333	202	3151.498	8.84E-05	0.9717

Table 3: Yoon Nelson and Thomas model parameters

Z (cm)	f (ml/min)	Thomas model parameters				Yoon Nelson model parameters			
		$K_{Th}$	$q_0$ (mg/g)	$q_{(exp)}$ (mg/g)	$R^2$	$K_{YN}$	$\tau$ (min)	$\tau_{(exp)}$ (min)	$R^2$
4.5	3	0.00055924	5.468	5.132	0.9916	0.0287	106.554	>97	0.9916
6	3	0.00034295	6.451	6.543	0.9725	0.0176	167.602	>165	0.9725
7.5	3	0.00017732	9.554	8.930	0.9386	0.0091	310.275	>290	0.9386
6	4.5	0.00040725	6.061	6.351	0.8422	0.0209	115.373	>110	0.8422

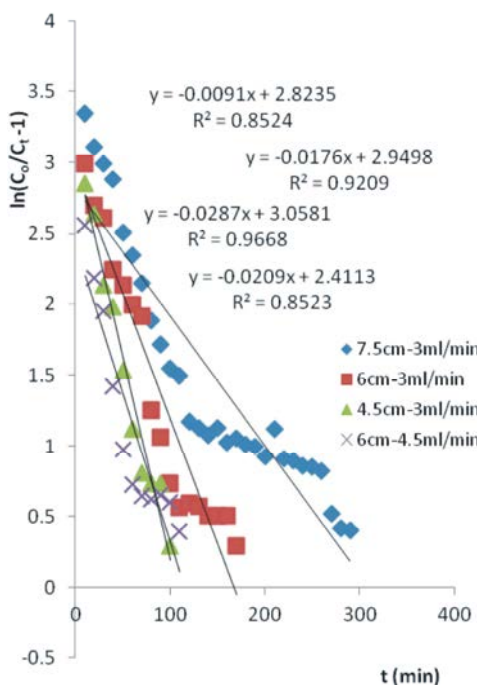


Fig. 5: Thomas model plot

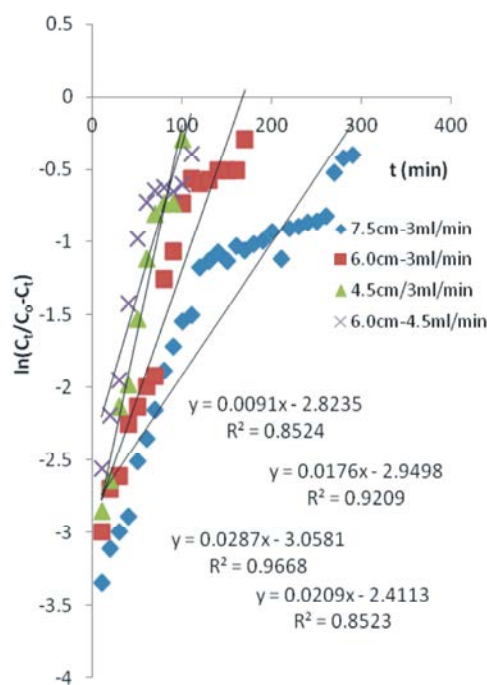


Fig. 6: Yoon Nelson model plot

**Column Models:** The Bed Depth Service Time (BDST) model parameters were obtained from Fig. 4 and are presented in Table 2. The large value of co-efficient of determination as observed in the table shows that the experimental data fit well in to the model. Therefore, BDST model could be employed in the adsorption of phenol onto rice husk activated carbon for the estimation of column design parameters such as bed depth and breakthrough time.

Thomas model parameters were obtained from the slope and intercept of Fig. 5 and are presented in Table 3. It was observed in the table that the theoretical breakthrough capacity ( $q_0$ ) increases as the bed depth increases and decreases as the flow rate increases.

This trend is in agreement with the earlier discussion on the effect of bed depth and flow rate on breakthrough capacity. Similar observation was reported by Han *et al.*, [28]; Noreen *et al.*, [36]. Hence, lower flow rate and higher bed depth favours better column performance. The table also shows that the theoretical breakthrough capacity obtained from Thomas model is close to the experimental breakthrough capacity indicating that Thomas model can be used to describe the column performance on the adsorption of phenol onto rice husk activated carbon.

Yoon Nelson model parameters were obtained from the slope and intercept of the plots in Fig. 6 and are presented in Table 3. The Table shows that the theoretical breakthrough time at 50%( $\tau$ ) increases with increase in

bed depth but decreases with increase in flow rate. This observation is in agreement with the earlier report on the effect of bed depth and flow rate on breakthrough time and similar trend was observed by Chowdhury *et al.*, [9]. The 50% theoretical breakthrough time obtained from Yoon Nelson model could be said to be very close to the experimental value as in each case, it was observed to be slightly more than the 40% experimental breakthrough time.

### CONCLUSIONS

The column experiment attained adsorption capacity of 28mg/g at breakthrough point of 0.5, flow rate of 4.5ml/min and bed depth of 7.5cm. As the column adsorption proceeds with time, it was observed that the adsorption capacity increases while the removal efficiency decreases. The performance of the column adsorption was significantly affected by bed depth and flow rate, thus improved column performance could be achieved by employing higher bed depth and/or lower flow rate. Analysis of the column experimental data indicates that models such as Thomas and Yoon Nelson can be employed in the estimation of design parameters and scale-up during pilot or full scale operations.

### REFERENCES

1. Ahalya, N., T.V. Ramachandra and R.D. Kanamadi, 2003. Biosorption of heavy metals. *Research Journal of Chemistry and Environment*, 7(4): 71-79.
2. Abdelkareem, M., 2013. Adsorption of Phenol from Industrial Wastewater Using Olive Mill Waste. *APCBEE Procedia*, 5: 349-357.
3. Girods, P., A. Dufour, V. Fierro, Y. Rogaume, C. Rogaume, A. Zoulalian and A. Celzard, 2009. Activated carbons prepared from wood particleboard wastes: Characterization and phenol adsorption capacities. *Journal of Hazardous Materials*, 166(1): 491-501.
4. Nabais, J.M.N., J.A. Gomes Suhas, P.J.M. Carrott, C. Laginhas and S. Roman, 2009. Phenol removal onto novel activated carbons made from lignocellulosic precursors: Influence of surface properties. *Journal of Hazardous Materials*, 167(1-3): 904-910.
5. Lazo-Cannata, J.C., A. Nieto-Márquez, A. Jacoby, A.L. Paredes-Doig, A. Romero, M.R. Sun-Kou and J.L. Valverde, 2011. Adsorption of phenol and nitrophenols by carbon nanospheres: Effect of pH and ionic strength. *Separation and Purification Technology*, 80(2): 217-224.
6. Benyahia, F., M. Abdulkarim, A. Embaby and M. Rao, 2006. Refinery Wastewater Treatment: A true Technological Challenge. The Seventh Annual UAE University Research Conference, UAE University, ENG. 186-193.
7. Kamble, S.P., P.A. Mangrulkar, A.K. Bansiwala and S.S. Rayalu, 2008. Adsorption of phenol and *o*-chlorophenol on surface altered fly ash based molecular sieves. *Chemical Engineering Journal*, 138(1-3): 73-83.
8. Bansal, M., U. Garg, D. Singh and V.K. Garg, 2009. Removal of Cr (VI) from aqueous solutions using pre-consumer processing agricultural waste: A case study of rice husk. *Journal of Hazardous Materials*, 162(1): 312-320.
9. Chowdhury, Z.Z., S.M. Zain, A.K. Rashid, R.F. Rafique and K. Khalid, 2013. Breakthrough curve analysis for column dynamics sorption of Mn (II) ions from wastewater by using *Mangostana garcinia* peel-based granular-activated carbon. *Journal of Chemistry*, Volume 2013, Article ID 959761
10. Lakshmi, U.R., V.C. Srivastava, I.D. Mall and D.H. Lataye, 2009. Rice husk ash as an effective adsorbent: Evaluation of adsorptive characteristics for Indigo Carmine dye. *Journal of Environmental Management*, 90(2): 710-720.
11. Ilaboya I., E. Oti, G. Ekoh and L. Umukoro, 2013. Performance of Activated Carbon from Cassava Peels for the Treatment of Effluent Wastewater. *Iranica Journal of Energy and Environment*, 4(4): 361-375.
12. Singh, K.P., A. Malik, S. Sinha and P. Ojha, 2008. Liquid-phase adsorption of phenols using activated carbons derived from agricultural waste material. *Journal of Hazardous Materials*, 150(3): 626-641.
13. Naiya, T.K., A.K. Bhattacharya, S. Mandal and S.K. Das, 2009. The sorption of lead (II) ions on rice husk ash. *Journal of Hazardous Materials*, 163(2-3): 1254-1264.
14. Krishnani, K.K., X. Meng, C. Christodoulatos and V.M. Boddu, 2008. Biosorption mechanism of nine different heavy metals onto biomatrix from rice husk. *Journal of Hazardous Materials*, 153(3): 1222-1234.
15. Mahalakshmi K., S. Suja, K. Yazhini, S. Mathiya and G.J. Kalaivani, 2014. A Novel Approach to Investigate Adsorption of Crystal Violet from Aqueous Solutions Using Peels of *Annona squamosa*. *Iranica Journal of Energy and Environment*, 5(2): 113-123.
16. Cadoni, P. and F. Angelucci, 2013. Analysis of incentives and disincentives for Rice in Nigeria. *Technical Notes Series*, MAFAP, FAO, Rome.

17. Kudaybergenov, K.K., E.K. Ongarbayev and Z.A. Mansurov, 2012. Thermally treated rice husks for petroleum adsorption. *International Journal of Biology and Chemistry*, 1: 3-12.
18. Dim, P.E. and J.O. Okafor, 2012. Development and characterization of adsorbent from coconut shells for purification of dye wastewater. *Journal of Science, Technology, Mathematics and Education*, 8(3): 68-80.
19. Amuda, O.S., A.O. Alade, A.O. Afolabi and F.E. Adelewo, 2012. Adsorption of Acenaphthene onto Activated Carbon Produced from Agricultural Wastes. *Journal of Environmental Science and Technology*, 5(4): 192-209.
20. Sugashini, S. and K.M. Meera Sheriffa Begum, 2013. Column Adsorption Studies for the Removal of Cr (VI) Ions by Ethylamine Modified Chitosan Carbonized Rice Husk Composite Beads with Modeling and Optimization. *Journal of Chemistry*, Volume 2013, Article ID 460971.
21. Dwivedi, C.P., J.N. Sahu, C.R. Mohanty, B.R. Mohana and B.C. Meikap, 2008. Column performance of granular activated carbon packed bed for Pb (II) removal. *Journal of Hazardous Materials*, 156(1-3): 596-603.
22. Sadaf, S. and H.N. Bhatti, 2014. Batch and fixed bed column studies for the removal of Indosol Yellow BG dye by peanut husk. *Journal of the Taiwan Institute of Chemical Engineers*, 45(2): 541-553. <http://dx.doi.org/10.1016/j.jtice.2013.05.004>.
23. Garg, U., M.P. Kaur, G.K. Jawa, D. Sud and V.K. Garg, 2008. Removal of cadmium (II) from aqueous solutions by adsorption on agricultural waste biomass. *Journal of Hazardous Materials*, 154(1-3): 1149-1157.
24. Saad, S.A., Isa, K. Md. and R. Bahari, 2010. Chemically modified sugarcane bagasse as a potentially low-cost biosorbent for dye removal. *Desalination*, 264(1-2): 123-128.
25. Lo, S.F., S.Y. Wang, M.J. Tsai and L.D. Lin, 2012. Adsorption capacity and removal efficiency of heavy metal ions by Moso and Ma bamboo activated carbons. *Chemical Engineering Research and Design*, 90(9): 1397-1406.
26. Kumar, U. and M. Bandyopadhyay, 2006. Fixed bed column study for Cd(II) removal from wastewater using treated rice husk. *Journal of Hazardous Materials*, 129(1-3): 253-259. doi:10.1016/j.jhazmat.2005.08.038.
27. Luo, X., Z. Deng, X. Lin and C. Zhang, 2011. Fixed-bed column study for Cu<sup>2+</sup> removal from solution using expanding rice husk. *Journal of Hazardous Materials*, 187(1-3): 182-189.
28. Han, R., Y. Wang, W. Yu, W. Zou, J. Shi and H. Liu, 2007. Biosorption of methylene blue from aqueous solution by rice husk in a fixed-bed column. *Journal of Hazardous Materials*, 141(3):713-718. doi:10.1016/j.jhazmat.2006.07.031.
29. Ozturk, N. and D. Kavak, 2005. Adsorption of boron from aqueous solutions using fly ash: Batch and column studies. *Journal of Hazardous Materials*, 127(1-3): 81-88.
30. Zhang, W., L. Dong, H. Yan, H. Li, Z. Jiang, X. Kan, H. Yang, A. Li and R. Cheng, 2011. Removal of methylene blue from aqueous solutions by straw based adsorbent in a fixed-bed column. *Chemical Engineering Journal*, 173(2): 429-436.
31. Ji, F., C. Li, J. Xu and P. Liu, 2013. Dynamic adsorption of Cu (II) from aqueous solution by zeolite/cellulose acetate blend fiber in fixed-bed. *Colloids and Surfaces A: Physicochem. Eng. Aspects*, 434: 88-94.
32. Kalavathy, H., B. Karthik and L.M. Miranda, 2010. Removal and recovery of Ni and Zn from aqueous solution using activated carbon from Hevea brasiliensis: Batch and column studies. *Colloids and Surfaces B: Biointerfaces*, 78(2): 291-302.
33. Song, J., W. Zou, Y. Bian, F. Su and R. Han, 2011. Adsorption characteristics of methylene blue by peanut husk in batch and column modes. *Desalination*, 265(1-3): 119-125.
34. Kundu, S. and A.K. Gupta, 2005. Analysis and modeling of fixed bed column operations on As (V) removal by adsorption onto iron oxide-coated cement (IOCC). *Journal of Colloid and Interface Science*, 290(1): 52-60.
35. Chen, N., Z. Zhang, C. Feng, M. Li, R. Chen and N. Sugiura, 2011. Investigations on the batch and fixed-bed column performance of fluoride adsorption by Kanuma mud. *Desalination*, 268(1-3): 76-82.
36. Noreen, S., H.N. Bhatti, S. Nausheen, S. Sadaf and M. Ashfaq, 2013. Batch and fixed bed adsorption study for the removal of Drimarine Black CL-B dye from aqueous solution using a lignocellulosic waste: A cost affective adsorbent. *Industrial Crops and Products*, 50: 568-579.
37. Mehdinia, S.M., K. Moeinian and T. Rastgoo, 2014. Rice husk silica adsorbent for removal of hexavalent chromium pollution from aquatic solutions. *Iranica Journal of Energy and Environment*, 5(2): 218-223.

#### Persian Abstract

**چکیده** صنایع همچون پالایشگاه مدام بمیزان بسیار زیادی فاضلاب حاوی غلظت بالا فنول دفع می کند. لذا چینیین فاضلاب بخاطر الاینده های موجود باید تصفیه شود. هدف اصلی این مقاله کاربرد کربن فعال از پوسته برنج است که در ستون جذب برای تصفیه فاضلاب پالایشگاه استفاده گردید. اثر بستر کربن فعال از پوسته برنج و دبی فاضلاب بر جذب فنول از فاضلاب مورد بررسی قرار گرفت. کلیاتی از بخشید. مدل های استفاده شده برای بستر ستون، زمان جذب، مدل های یان و توماس با استفاده از داده های تجربی مورد انالیز قرار گرفت. آزمایش ظرفیت جذب ۲۸ mg/g دبی ۴/۵ ml/min و عمق بستر ستون ۷/۵ cm و نقطه رخنه ۰/۵ حاصل گردید. کارایی ستون جذب متاثر از دبی جریان و عمق بستر ستون بوده و مدل های یان و توماس برای تخمین پارامترهای ستون جذب در مقیاس صنعتی مورد استفاده قرار گرفت.