



## Evaluation of operative factors on conversion efficiency of biodiesel production from waste cooking oil

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### A B S T R A C T

Today, the production of biodiesel from waste cooking oil due to its renewability, low cost of production and also low pollution become more popular. In this research, the factors affecting the production of biodiesel by trans-esterification method from waste cooking oil were evaluated. For this purpose, a batch reactor which equipped with a mechanical stirrer, was designed and fabricated. Waste cooking oil as feedstock and methanol were used with molar ratio of 1:6 in the presence of potassium hydroxide as a catalyst (1wt%). In order to optimize the production process, these three parameters including rate of stirring at 3 levels (450, 550 and 650 rpm), at four temperatures 40, 50, 60 and 70°C, reaction time at 6 levels (10, 20, 30, 40, 50, 60 min) were considered. The results indicated, more than 90% of ester conversion occurred in the first 30 minutes and biodiesel production increased and reached to its maximum amount at stirring speed of 650 rpm and also the highest rate of esters conversion occurred at 60 °C. Thus, 60°C for reaction temperature and 650 rpm for stirring speed at 30 min were optimum production condition in order to reach to the most amount of biodiesel from waste cooking oil using a batch stirred tank reactor.

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## INTRODUCTION

In general, renewable and nonrenewable energy resources can be divided into two categories. Renewable energy known as a sustainable fuel sources and often include solar, wind, biomass, biofuels, tidal, etc [1]. Renewable-energy development has become as important part of the global energy policy to reduce the greenhouse effect caused by fossil fuels [2]. Due to the geographical extent of the scattering centers and remote energy consumption, use of renewable energy such as hydropower, wind, solar, geothermal, biomass, bio-fuels, biogas, is more economical.

Biodiesel known as an environmental friendly fuel to produce energy approximately equal to the diesel fuel. Hence it can be considered as the most efficient alternative to diesel fuel. Biodiesel considered as an approximately free of sulfur fuel (sulfur-free <10 ppm) and a total reduction of pollutants such as carbon dioxide, unburned hydrocarbons and smoke is visible in comparison with diesel fuel [3].

Biodiesel known as an alternative fuel with a high flash point, with low emissions and high security that can be used in existing diesel engines and equipment used

atomization without any reduction in engine performance [4].

In the most recent researches on biodiesel production from vegetable oils in batch reactors with mechanical stirring, temperatures near the boiling point of the alcohol was used [5,6,7]. The reactor tank is filled with waste oil and subsequent alcohol and a catalyst for the reaction of trans-esterification processes can be added. After stirring completed, the contents of the reactor transmitted to water washing tank in order to wash [8].

The catalyst used for the production of biodiesel may be alkaline, acid or enzymes. Potassium hydroxide, sodium hydroxide and sodium methoxide are the alkaline catalyst used for the process [9]. Based on research conducted using homogeneous catalysts such as potassium hydroxide and sodium hydroxide; the process is speed up reaction time in comparison to sulfuric acid and phosphoric acid which are acid catalysts. It is noted that the acid esterification reaction used as a pretreatment for raw materials with high water content and fatty acid [10,11,12].

Various stirrer has been used to improve reaction and reduce biodiesel production time [13], including mechanical stirrer, magnetic stirrer [14] turbines [15], anchor stirrer [16], Helical stirrer [17], etc. Most researchers in the field of mechanical stirring focused on

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type of stirrer. Their report suggests that the mechanical agitator capable of producing a product with a purity of more than 97% [18].

Meng et al. [19] concluded that the use of batch reactor for the production of biodiesel from jatropha oil, 93% of triglycerides were converted into methyl esters. Batch stirred tank reactor was operated with molarity of alcohol to triglyceride ratio ranging from 4: 1 to 20: 1 and the operating temperature range is reported 25-85°C and at temperatures above the boiling point of alcohol, reactor should be equipped with a condenser [20].

The objectives of this research are to find the suitable time, temperature and optimized condition to produce biodiesel from waste cooking oil using a batch reactor equipped with mechanical stirrer.

## MATERIALS AND METHODS

### Specification

Figure 1 shows reactor system which designed (Solidworks 2012) and fabricated, different parts of reactor are demonstrated in Figure 1. The device is designed for a capacity of 30 liters of liquid in the tank.

To facilitate phase separation during discharge and also ensure a uniform mixing bottom part of the tank was built in the shape of a cone. Reaction and leaching tanks were built separately; heating system parts includes (1000-watt heater, temperature sensors and digital display of temperature) for heating the liquid and keep the temperature at a constant value. The signal from the sensor turn on or off the power to the heater.

Power Transmission system (electric motor (1.5 KW), inverter (hi run model N100 for measuring the rotational speed of the stirrer), Electric pump (0.37 KW) were used for mixing and transport raw materials (reactants) and move products between the tanks. The reactor frameworks was made with a 3 × 5 cm profile. The bearings were used for each axis to prevent horizontal and vertical vibrations [21].

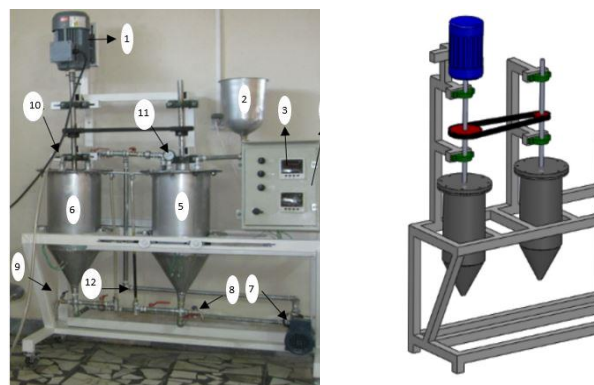
### Test method

To perform the test, molarity of 6: 1 oil to alcohol and KOH catalyst was 1% of oil by weight selected. Tests based on three levels of the factor variable stirrer speed (650, 550, 450 rpm), and the reaction temperature at four levels (40, 50, 60, 70°C) and a reaction time of 6 levels (10, 20, 30, 40, 50, 60 min) was performed. The waste cooking oil at a rate of 8 liters per test batch was chosen.

Two factors, free fatty acids and water, affected the reaction procedure to saponification, and thus reduce the quality and quantity of product. As a result of oil with a high content of free fatty acid (more than 0.05%) requires pre-treatment stage (stage esterification).

After the reaction is finished, liquid divided in two parts, the upper layer is biodiesel and bottom layer is glycerin. In accordance with ASTM, biodiesel glycerin

content should be less than 0.24% of produced glycerin. Separation of the products also depends on the quality of oil and the soap made in reaction, with time needed to



**Figure 1.** Actual experimental rig with designed and fabricated reactor (1: Electrical motor, 2: Input of reaction tank, 3: Digital display of temperature, 4: electrical panel, 5: reaction tank, 6: water washing tank, 7: electrical pump, 8: Output of reaction tank, 9: Output of water washing tank, 10: input of water washing tank, 11: pressure gauge, 12: flush hose)

settle down will vary.

## RESULTS AND DISCUSSION

Biodiesel fuel properties according to ASTM standard and measured values are provided in Table (1). According to the results, the characteristics of biodiesel match within the standard ASTM D6751.

**TABLE 1.** Characteristics of Biodiesel Standards and Measurement

Characteristics	Unit	ASTM Standard	Measurement
Flash point	C°	Min 93	172
Water content	% volume	Max 0.050	0.026
Viscosity(at 40 C°)	mm <sup>2</sup> /s	1.9-6	3.4
Cloud point	C°	-	-2
Carbon residue	% mass	Max 0.050	0.042
Acid number	mg KOH/gr	Max 0.5	0.321
Free glycerine	% mass	Max 0.024	0.007

Analysis of main variance and interaction affected on temperature, time and stirring, which is provided in Table (2). The results showed that stirring has significant effect on efficiency of conversion of free fatty acids at 1wt% level. The effect of time on biodiesel conversion efficiency is significant at 1% level. Effect of transesterification reaction temperature on the biodiesel conversion efficiency is significant at 1% level.

**TABLE 2.** Analysis of variance factors affecting conversion efficiency of biodiesel

Variable	Mean square	Degrees of freedom	Conversion efficiency	F-value
Stirring rate (V)		2	30832.7	41110.22 **
Reaction time (t)		5	13839.5	18452.62 **
(V*T)		10	110.3	147.02 **
Temperature (T)		3	7397.5	9862.33 **
(V*T)		6	0.074	0.0989 ns
(t*T)		15	0.026	0.0346 ns
(V*t*T)		30	0.019	0.0256 ns
Error		144	0.75	
The coefficient of variation (CV)			10.11	

\*\* and ns are significant and non-significant at the 1% level, respectively.

The results suggest that the interaction between the stirring speed and reaction time ( $V \times t$ ), the conversion efficiency of biodiesel at 1% level is significant. If the interaction of temperature and stirring in the study area as well as the interaction between time and temperature on the conversion efficiency of biodiesel is not meaningful. This means that in any case using the excessive heat (over  $^{\circ}\text{C}$  60) evaporates the alcohol (methanol) which stops the reaction and will not affect on conversion efficiency of biodiesel.

### The effect of reaction temperature

The reaction temperature is an important factor in the production of biodiesel. The results in Table 3 are given the reaction temperature increased from 40 to  $60^{\circ}\text{C}$ , the efficiency of biodiesel production has increased from 72.14 to 95.81%. Increasing the temperature increases the reaction rate and mass transfer in the reaction tank. With further increase of the temperature to  $70^{\circ}\text{C}$ , biodiesel conversion efficiency decreased; that could be due to evaporation of alcohol (close to the boiling point of methanol). Knothe [22] reported more than 95% conversion reached at  $60^{\circ}\text{C}$  and reaction time was 3 hours.

**TABLE 3.** Average effect of reaction temperature on the conversion efficiency of biodiesel

Temperature, $^{\circ}\text{C}$	Biodiesel conversion efficiency %
40	72.14 c
50	81.36 b
60	95.81 a
70	94.92 a

The optimum temperature of  $80^{\circ}\text{C}$  with a performance of 80% was concluded [22]. Koria and Thangaraj [23] at the best temperature for the production of biodiesel obtained  $60\text{-}65^{\circ}\text{C}$ . According to the results obtained in this study, 95.81% efficiency at a lower temperature than the boiling point of methanol ( $69^{\circ}\text{C}$ )  $60^{\circ}\text{C}$  was obtained, which saves time and energy. As a

result, the reaction temperature of  $60^{\circ}\text{C}$  was the optimum reaction temperature for the production of biodiesel from waste vegetable oil in a stirred tank batch reactor. Please note that the use of a mechanical stirrer makes better heat transfer between the reactants.

### The effect of reaction time

Table (4) shows the effect of conversion efficiency of biodiesel, in a reaction time in the range of 10 to 60 minutes. The results showed that biodiesel conversion efficiency gradually increased from 43.3% to 95.3% in the reaction time of 10 to 30 minutes. When the stirring process starts the Trans esterification reaction begins between methanol and oil in the presence of potassium hydroxide.

The response rate was primarily due to the low dissolution of alcohol and oil, and second, by dissolving those increases and finally reduced consumption of reactants and reaches to equilibrium. In addition, the response time greater than 30 minutes, made no change in the conversion efficiency of the reactor. That was probably due to the ability of mixing stirrer reactor, which is designed to reduce the reaction time to 30 minutes.

Efficiency of 95% biodiesel production using potassium hydroxide catalyst with an alcohol such as methanol, ethanol and 1-butanol have been reported in literature [24]. The effect of reaction time on the yield of biodiesel production using the catalyst of calcium hydroxide at a temperature  $60^{\circ}\text{C}$  and the speed of the stirring 300 rpm and the molar ratio of methanol to oil 15: 1 and the amount of catalyst to 6% by weight of waste vegetable oil were experimented. The results found that trans esterification reaction time is dependent on the severity of the reaction condition [25].

**TABLE 4.** effect Average of reaction time on the conversion efficiency of biodiesel

Reaction time (min)	Biodiesel conversion efficiency average
10	43.3 c
20	77.3 b
30	95.3 a
40	94.3 a
50	94.0 a
60	94.7 a

By changing the reaction time in the range of 0 to 120 minutes, biodiesel production efficiency significantly increased [26]. Decrease in efficiency was shown that can be caused by operation at low temperature which was caused by increasing oil viscosity due to low temperature. Reduction of two-phase separation of methanol and oil mixture. The optimum reaction time for a batch reactor was set at 30 minutes.

### The effect of mixing speed

Table 5 shows the effect of stirring velocity on the conversion efficiency of biodiesel, biodiesel production increases with increasing stirrer speed and reaches its maximum at the speed 650rpm. With increasing stirring rate, collisions between molecules which are insoluble in alcohol and oils happened and constitute a heterogeneous phase and more mass transfer occurs.

**TABLE 5.** The effect of stirring rate on biodiesel conversion efficiency

Conversion efficiency	Stirring rate (rpm)
42.3 c	450
83.5 b	550
94.5 a	650

Mean that shown with common letters are not significantly different.

High stirring rate, leads to high conversion rate and reaction time will be lower. However, excessive stirring rate may reduce the conversion rate of the reaction [28]. Hoseini [28] found that more than 90% biodiesel conversion happened in a the reactor with stirrer speed of 900 rpm, it was also stated that further increase in mixing speed does not affect on the progress of the reaction .

### CONCLUSION

In this study, biodiesel was produced from waste cooking oil using a batch stirred tank reactor which designed and constructed. Results showed that the efficiency of conversion of waste cooking oil into biodiesel using stirred tank reactor at optimum temperature 60°C, the stirring speed of 650 rpm and the reaction time of 30 minutes using methanol as alcohol and potassium hydroxide as a catalyst.

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**Persian Abstract**

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**چکیده**

امروزه بیودیزل تولیدی از روغن پسماند به علت آلاینده‌گی و هزینه تولید کمتر و تجدید پذیری، بیش از پیش مورد توجه قرار گرفته است. در این تحقیق عوامل موثر بر تولید بیودیزل به روش ترنس-استیریفیکاسیون از روغن پسماند آشپزی مورد بررسی قرار گرفت. برای این منظور یک راکتور همزن-دار بیج، طراحی و ساخته شد. روغن پسماند آشپزی به عنوان ماده اولیه و متانول به عنوان الکل با نسبت مولی ۱:۶ در حضور پتاسیم هیدروکسید به عنوان کاتالیزور به میزان ۱٪ وزنی روغن پسماند استفاده شد. بهینه سازی فرآیند تولید، در سه پارامتر دور همزنی در ۳ سطح (650 rpm, 550, 450)، دما در ۴ سطح (40°C, 50°C, 60°C, 70°C) و زمان واکنش در ۶ سطح (۶۰، ۵۰، ۴۰، ۳۰، ۲۰، ۱۰ دقیقه) انجام شد. نتایج نشان داد، تبدیل بیش از ۹۰٪ استرها در ۳۰ دقیقه اول اتفاق افتاد. در دور همزنی 650 rpm میزان تولید بیودیزل افزایش یافته و به حداکثر میزان خود رسید. بیشترین میزان تبدیل استرها در دمای 60°C اتفاق افتاده است. بنابراین زمان واکنش ۳۰ دقیقه و دور 650 rpm و دمای واکنش 60°C بهینه-ترین حالت برای تولید بیودیزل از روغن پسماند با استفاده از راکتور همزن-دار بیج می-باشد.

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